Intermolecular Attractions (v3)

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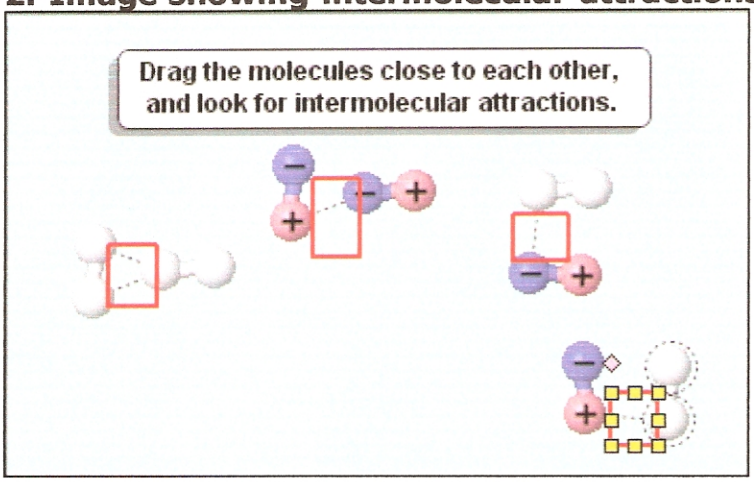
Class: Chem A, Period C - 09-10
Other Group Members:

1. What is the difference between polar and non-polar molecules? (Be sure to include more than the colors used to represent them in your response!)

In the non-polar molecules there is an uneven surface charge that is constantly changing. In addition the bonds bewtween the atoms are not constant. The polar molecules' surface charge is constantly eaven and there is always constant bonds between teh atoms.

2. Image showing intermolecular attractions:



3. Which of the following formed intermolecular attractions (check all that apply):

- + ends of polar molecules to - ends of other polar molecules
- non-polar molecules to other non-polar molecules
- non-polar molecules to + parts of polar molecules
- non-polar molecules to - parts of polar molecules

4. What is the primary attraction between NON-POLAR molecules:

- London Dispersion attraction

5. What is the primary attraction between POLAR molecules:

- dipole-dipole attraction

6. Which type of intermolecular attraction is strongest:

- dipole-dipole attraction